



Facile synthesis of a modified HF-free MIL-101(Cr) nanoadsorbent for extraction nickel in water and wastewater samples

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ABSTRACT

A novel sorbent based on MIL-101(Cr) nanoadsorbent as a MOF structure was used for nickel extraction from water and wastewater samples. In this study, 30 mg of MIL-101(Cr) nanoadsorbent dispersed in 50 mL of water or wastewater samples, after sonication and adjusting pH =8.5, the nickel ions was extracted by carboxyl groups of terephthalic acid ($\text{MOF}-(\text{C}_6\text{H}_4 \text{ (COO)}^2-\dots \text{ Ni}^{2+})$) by dispersive suspension-micro solid-phase extraction (DS- μ -SPE). The MOF was separated from liquid phase with filter membrane (0.2 μm), eluted with 0.5 mL of nitric acid as back-extraction solution and finally, the nickel concentration in elution was determined by atom trap-flame atomic absorption spectrometry (AT-FAAS) after dilution with DW up to 1 mL. The LOD, the linear range and preconcentration factor were achieved $1.5 \mu\text{g L}^{-1}$, $5-160 \mu\text{g L}^{-1}$ and 49.7, respectively. The absorption capacity of MOF for nickel was obtained 136.8 mg g^{-1} . The results of procedure were validated by spiking of samples and ET-AAS analyzer.

1. Introduction

Nickel is the 28th element of Periodic table that found in abundance the earth's crust [1-4]. Pure nickel is a hard, silvery white, lustrous, malleable, ductile, shining metal with high electrical and thermal conductivity [2,5]. The soluble nickel salts and metallic nickel, nickel sulphides and nickel oxides are poorly water soluble [2]. Nickel as a heavy metal, nonessential great environmental concern because widely occurring in the environment from various natural sources including emissions from fossil fuel consumption and anthropogenic processes

including modern technologies and production of products such as coins, jewelry, stainless steel, batteries, medical devices, plating and welding [6-9]. In 2008, nickel received the shameful name of the "Allergen of the Year" [8]. Humans are exposed to nickel in both occupational and Non-occupational forms. Occupational exposure to Nickel is primarily associated with workers in the producing and using of Nickel in industry sectors. Non-occupational sources of nickel exposure include is food, air and water. Human exposure to nickel occurs of inhalation, dermal contact, and gastrointestinal primary routes. Some specific aspects of nickel toxicities include Genotoxicity, Developmental toxicity, Neurotoxicity, Haematotoxicity, Immunotoxicity, Neurotoxicity

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and Carcinogenicity is [2, 4, 5, 8-11]. The adverse health effects of nickel for humans depend on the route of exposure, water solubility of nickel compounds, dose, bodyweight, sensitivity and exposure periods [10]. Exposure to nickel causes irritation of the nose, vertigo, insomnia, sinuses and loss of sense of smell, headache, nausea, vomiting, chest pain, nonproductive cough, dyspnoea, cyanosis, abdominal pain, diarrhea, tachycardia, palpitations, sweating, visual disturbances, weakness, lassitude and shortness of breath, giddiness and it could also lead to asthma, bronchitis and other respiratory diseases, eventually causing lung cancer [4,8,12]. According to the Research on Cancer (IARC) and the U.S. Department of Health nickel sulfate, sulfides and oxides combinations of nickel are be classified in Group 1, (i.e. cancerogenic to humans) and metallic nickel in Group 2B, (i.e. Possibly carcinogenic to humans [2,5,8,13]. The Permissible Exposure Limit (PEL) by Occupational Safety and Health Administration (OSHA) for dust, fume and metal of nickel and Recommended Exposure Limit (REL) by National Institute of Occupational Safety and Health (NIOSH) for metal and soluble of nickel are 1and 0.015 mg m³ respectively. In addition, World Health Organization (WHO) has proposed a guideline value of 20 µg L⁻¹ for the maximum permissible concentration of nickel in drinking water [14, 15]. A variety of methods such as cloud point extraction (CPE) [16-20], ionic liquid dispersive liquid-liquid micro extraction (IL-DLLME) [18,21], liquid-liquid extraction, cold induced aggregation microextraction (CIAME) [15], Solid phase extraction (SPE) [22-26], ligandless-ultrasound-assisted emulsification microextraction (USAEME) [27], coupled to various instrumental techniques like flame atomic absorption spectrometry (FAAS) [16,23,28,29,30-32] graphite furnace atomic absorption spectrometry (GFAAS) [33-34], ultrasonic nebulizer and inductively coupled plasma optic emission spectrometry (USN-ICP-OES) [17], inductively coupled plasma atomic

emission (ICP AES) [35], UV spectrometry [36], X-ray fluorescence spectrometry [37], inductively coupled plasma optic emission spectrometry (ICP-OES) [21], high performance liquid chromatography (HPLC) [38], inductively coupled plasma optic emission spectrometry (ICP-OES), have been developed for the for preconcentration and determination of nickel at low concentrations [39]. During the last decade, metal-organic frameworks (MOFs) have widely attracted international attention due to their high thermal stability, large surface area, and pore volume. Although a variety of MOFs has been developed, having their own particular properties, MIL (standing for materials of institute Lavoisier) structures have been more intriguing than other constituents owing to their extra-large cavities, high moisture stability, and lower production cost. For the first time, Ferey and his colleagues succeeded to synthesize nanoporous chromium terephthalate MIL-101(Cr) with a 3D structure by incorporating ter-ephthalic acid (benzene-1,4-dicarboxylic acid), chromium salt (Cr (NO₃)₃·9H₂O) and hydrofluoric acid (HF), as a modulator, in an aqueous medium [40]. In addition to the substantially large surface area and pore volume, they displayed that MIL-101(Cr) has an exclusive pore size distribution with two types of inner cages (29 Å and 34 Å) and windows (12*12 Å and 14.7*16 Å). These promising features have unprecedentedly boosted the role of MIL-101(Cr) and its derivatives in the gas storage (CO₂ [43-44], CH₄ [45-49] and H₂ [50-53]), gas separation (CO₂/CH₄ [54-56] and CO₂/N₂ [57-59]) and catalytic applications [60-63].

2. Material and Methods

2.1. Instrumental

The nickel concentration was determination by spectra GBC 906 double beam atom trap flame atomic absorption spectrophotometer with deuterium lamp as a background correction (AT-F AAS, GBC, Aus). The atom trap was installed on an air-acetylene burner. The operating software of

AVANTA was utilized for collecting and storing data. A HCL of nickel adjusted at a current of 4.0 mA and a wavelength of 232.0 nm with a spectral bandwidth of 0.2 nm. The working range for nickel was obtained 0.5 -8 mg L⁻¹. The pH-meter (Metrohm 744 Switzerland), centrifuge (EBA, Germany) and ultrasonic bath (Kunshan) were used.

2.2. Reagents and Materials

In the synthesis procedure, chemicals including Chromium nitrate nonahydrate ($\text{Cr}(\text{NO}_3)_3 \cdot 9\text{H}_2\text{O}$, 97%), 1, 4-benzene dicarboxylic acid (H_2BDC), Methanol (MeOH, Merck, 99.9%), ethanol (EtOH, Merck, 99.9%), N, N-dimethylformamide (DMF, Merck, 99.8%) and Acetone were purchased from commercial vendors and utilized as received. (DI) water was employed as solvents. All reagents with analytical grade purchased from Merck (Darmstadt, Germany). Stock solutions of Ni(II) were prepared by dissolving powder amounts of $\text{Ni}(\text{NO}_3)_2$ in deionized water (DW). The standard solutions were prepared daily by diluting with deionized water. Deionized water prepared from Milli-Q plus water from Millipore, USA. The pH of solutions was adjusted by ammonium chloride ($\text{NH}_3/\text{NH}_4\text{Cl}$) for pH 8–10. All the laboratory glasses were cleaned by 10% (v/v) nitric acid for 24 h and washed with DW 10 times.

2.3. Preparation of MIL-101(Cr) nanoparticles

In a typical procedure, 16 g of $\text{Cr}(\text{NO}_3)_3 \cdot 9\text{H}_2\text{O}$ and 6.56 g of terephthalic acid were separately added to deionized water (200 mL) and mixture was intensively stirred for 30 min with a magnetic stirrer. The resulting mixture was transferred to a stainless steel autoclave and heated at 220 °C for 18 h. Upon heating, the mixture became as a soft green powder which was washed five times with boiling deionized water, three times with MeOH and three times with acetone in order to remove impurities and unreacted materials. In addition, to separate the remained terephthalic acid, the resulting solid was suspended in 50 mL of dimethylformamide and kept at 70°C overnight. After cooling, the

sample was washed by pure ethanol for three times and the resulting powder was dried at 100 °C for 24 h to obtain the final product.

2.4. Characterization

To investigate the crystallinity and phase structure of the samples, powder X-ray diffraction (PXRD) method was considered by using Philips PW-1730 instrument with Cu-K α radiation ($\lambda = 1.5406 \text{ \AA}$). The scanning rate was 1 deg/min and operating power was 40 kV and 40 mA. To determine the functional groups, Fourier transform infrared spectroscopy (FT-IR) was performed using Bruker (VERTEX 70) spectrum from 400 to 4000 cm⁻¹. Field Emission Scanning Electron Microscopy (FE-SEM) studies were carried out with a JEOL JEM 3010 instrument under 15 kV voltage and 50 kx and 100 kx magnifications. The specific surface area, total pore volume, average pore width and pore size distribution measurements was performed with a Micrometrics ASAP-2010 instrument by adsorption of nitrogen at 77 K. Before analysis, the sample was degassed at 175°C and vacuum pressure for 2 h to remove moisture, solvents and other unwelcome molecules from pores.

2.5. Analytical Procedure

By DS-μ-SPE procedure, the nickel ions were separated/preconcentrated from 50 mL of wastewater based on MIL-101(Cr) as MOF nanoadsorbent. First, the pH of wastewater samples and standard nickel solution containing 1–400 µg L⁻¹ was adjusted up to 8.5 with phosphate or ammonium chloride buffer before adding 30 mg of MIL-101(Cr). After shaking in ultrasonic bath for 5 min at room temperature (50 kHz, 100 W), the nickel ions was extracted by carboxyl groups of MIL-101(Cr). Then the MIL-101 sorbent was separated from liquid phase with filter membrane (0.2 µm) based on vacuum accessory and eluted with 0.5 mL of nitric acid as back-extraction solution. Finally, the nickel concentration in eluent was determined by atom trap-flame atomic absorption spectrometry (AT-FAAS) after dilution with DW up to 1 mL (Fig. 1).



Fig.1. Nickel extraction based on MIL-101(Cr) from wastewater samples by DS- μ -SPE procedure

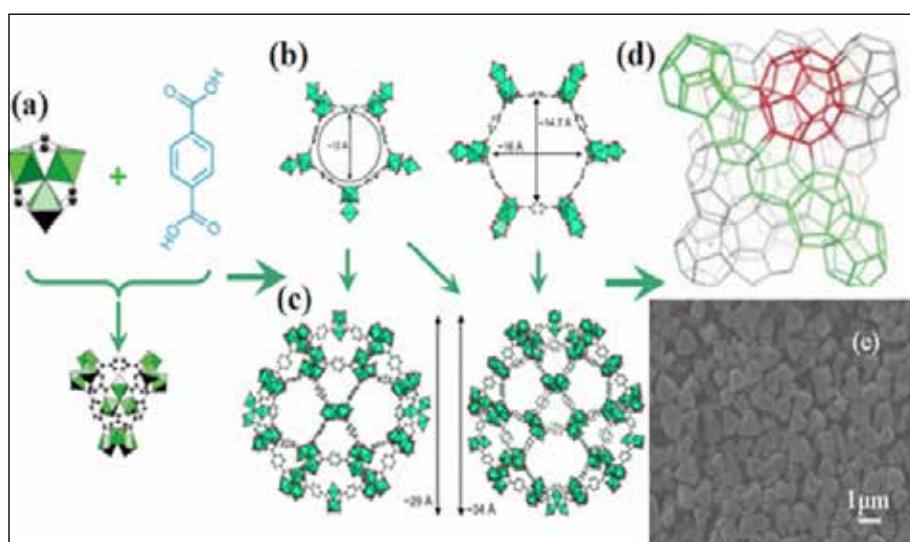


Fig.2. The robust framework of MIL-101 based on chromium (III) octahedral clusters

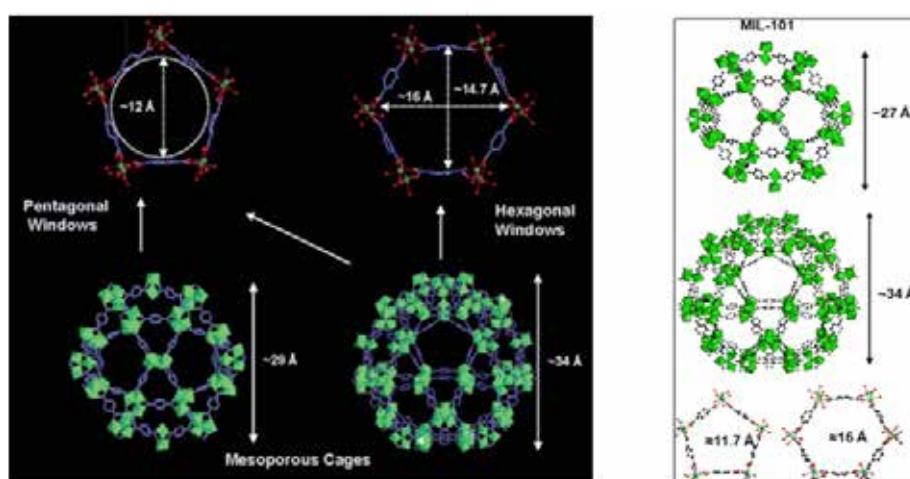


Fig.3. The structure of two types of mesoporous cages in MIL-101 with pentagonal forms

3. Results and discussion

Férey et al. prepared MIL-101 (MIL, Mat'erial Institut Lavoisier) with a chemical composition of $\{Cr_3F(H_2O)_2O(BDC)_3.nH_2O\}$ ($n \approx 25$; 1, 4-benzenedicarboxylate (BDC) and superior physicochemical properties. The robust framework of MIL-101 was comprised of trimeric chromium (III) octahedral clusters interconnected by BDC molecules resulting in an augmented MTN zeotype structure (Fig. 2).

This structure was comprised of two types of mesoporous cages with diameters of ~ 29 and 34 \AA accessible through two types of microporous windows (the smaller cages have pentagonal windows with a free opening of $\sim 12 \text{ \AA}$, while the larger cages possess both pentagonal and hexagonal windows with a $\sim 14.7 \text{ \AA}$ by 16 \AA free aperture) (Fig. 3).

3.1. Extraction Mechanism

The material exhibits excellent stability against moisture and other chemicals, and the terminal water molecules in MIL-101 can be removed by

heating in air or under vacuum at 423 K , which generates two coordinatively unsaturated open metal sites (CUS) per trimeric Cr(III) octahedral cluster. In addition to its highly porous nature, MIL-101 has attracted considerable attention because functional modifications on MIL-101 can be achieved easily either directly using a functionalized ligand during the synthesis or indirectly via the diverse post-synthesis chemical treatment on the CUS or on the organic linkers. Among the MOFs known, MIL-101 is one of the most promising porous materials for future energy and environmental applications, surpassing MOF-5 or HKUST-1, owing to its superior physicochemical properties including high hydrothermal/chemical stability and desirable textural properties. Due to MIL-101 Structure nickel was physically and chemically extracted based on porous materials in MIL-101 nanostructure and covalence bonding by carboxyl groups of terephthalic acid (MOF-($C_6H_4(COO)^2$) $\dots Ni^{2+}$), respectively (Fig. 4).

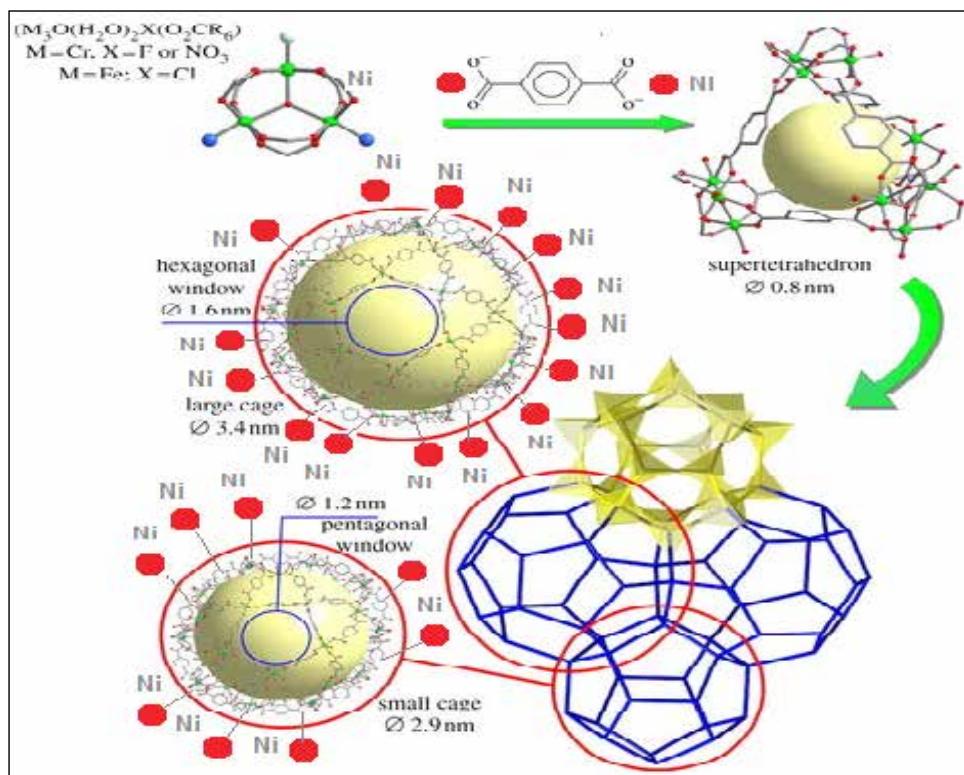
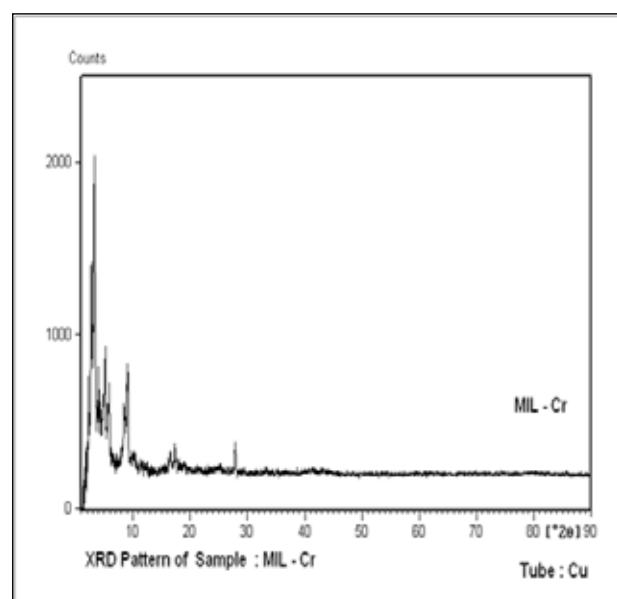


Fig. 4. The extraction mechanism of nickel by MIL-101 nanostructure

3.2. PXRD patterns

PXRD pattern of the MIL-101(Cr) sample, prepared without modulators, is depicted in Figure 5. As it stated, MIL-101(Cr) possessed the most intensive peaks, as an indicator of the most crystallinity. The negligible peaks at around $2\theta = 17.4^\circ$, 25.2° or 27.9° affirmed successful removal of unreacted H₂BDC crystals from the sample framework [36]. Moreover, the main diffraction peaks of MIL-101(Cr) ($2\theta \approx 5.25^\circ$, 8.55° , 9.15° , and 16.58°) was thoroughly compatible with the MIL-101-HF-1, as a general reference for MIL-101(Cr).



3.3. FE-SEM

The morphology of the prepared MIL-101(Cr) sample was characterized by field emission scanning electron microscope (FE-SEM) analysis. As it is seen in Figure 6, the image of MIL-101(Cr) shows highly crystalline octahedral morphology for this material which was compatible with the reported MIL-101(Cr) structures. In addition, there were no needle-shaped crystals in the images representing the complete removal of H₂BDC crystals by the post-purification process.

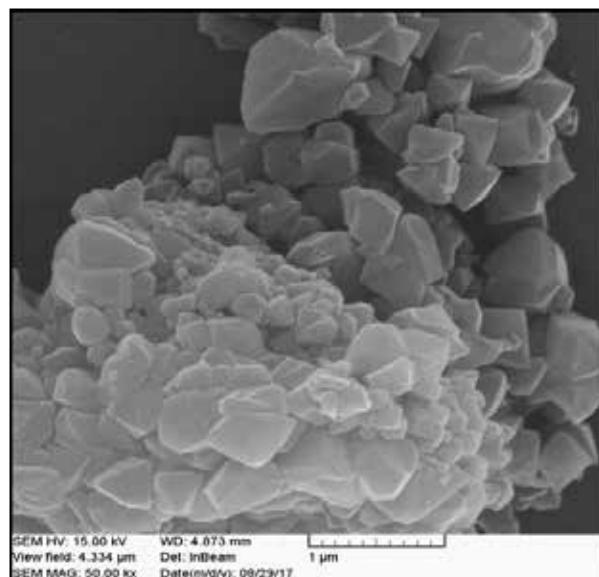
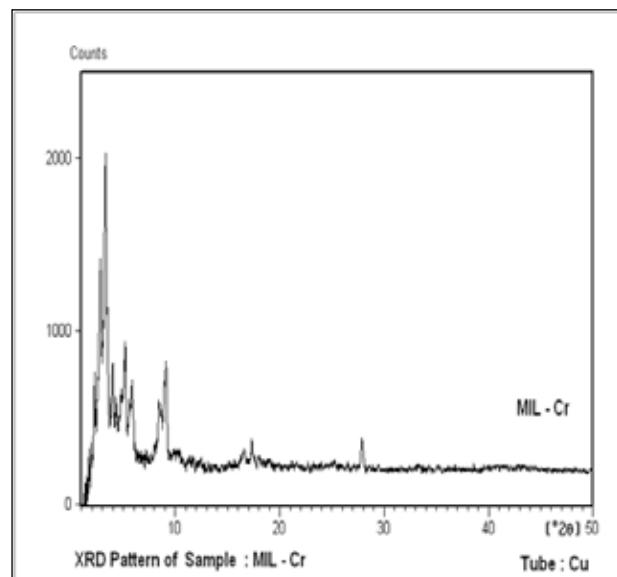
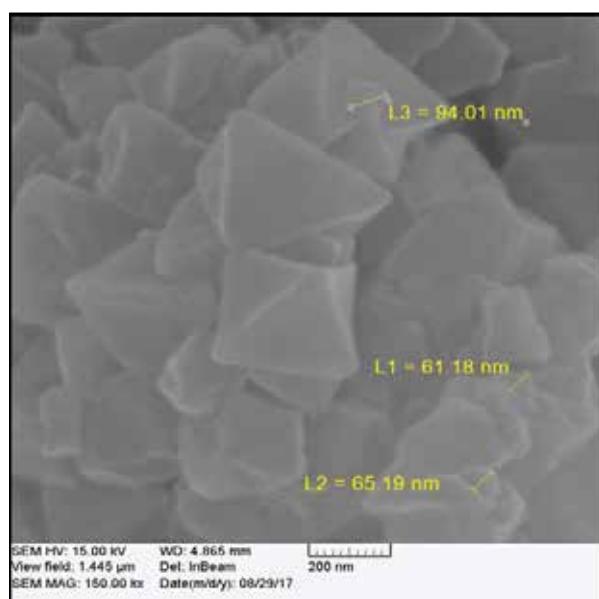


Fig. 5. The PXRD pattern of the MIL-101(Cr) sample

Fig. 6. The morphology of MIL-101(Cr) by field emission scanning electron microscope (FE-SEM)

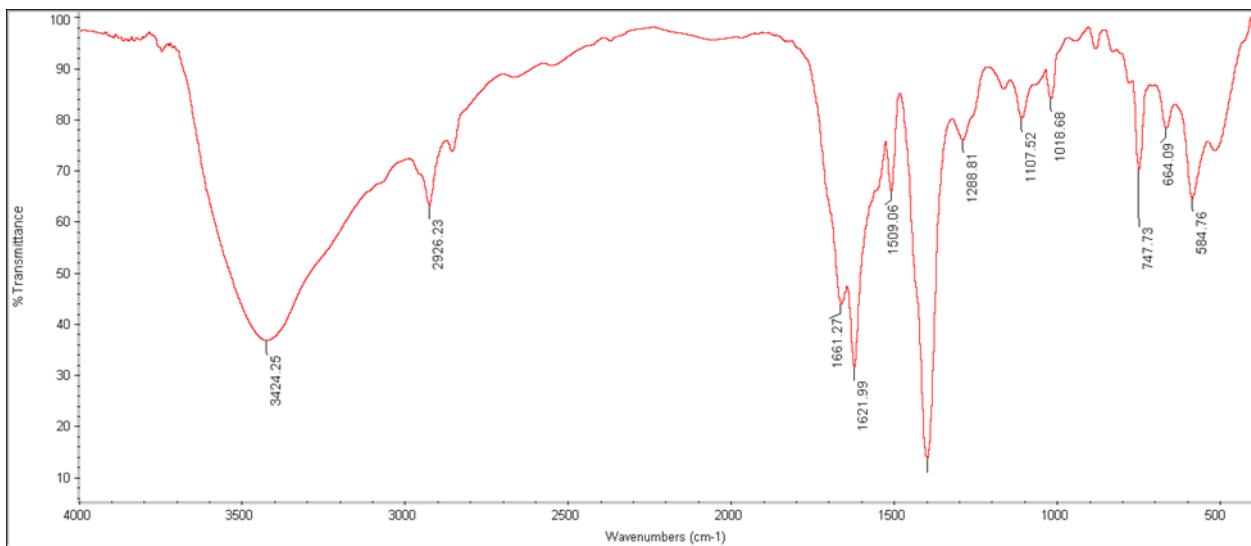


Fig. 7. FT-IR spectra of the synthesized MIL-101(Cr)

3.4. FT-IR

FT-IR spectra of the synthesized MIL-101(Cr) illustrated in Figure 7. In this spectrum the peak around of 570 cm^{-1} was assigned to the Cr–O stretching vibration, reflecting the formation of MIL-101(Cr) structure. The peaks of between 600 and 1600 cm^{-1} were indicated to H₂BDC and its aromatic rings. The bands at $750, 884, 1160\text{ cm}^{-1}$ were attributed to the C–H bond in CH₃ group and the peak observed at 1508 cm^{-1} indicates the C=C stretching. A strong band at 1404 cm^{-1} is related to O–C–O symmetric vibrations and showing the dicarboxylate moiety in the sample. The typical bands located at 1625 and 3400 cm^{-1} confirmed the presence of hydroxyl groups or moisture in the sample. The FT-IR results were in accordance to the previous reports FT-IR patterns which could be an evidence for the formation of MIL-101(Cr) structures (Fig. 7).

3.5. Physical properties

To determine physical properties of synthesized MIL-101(Cr) sample, the N₂ sorption test at 77 K was carried out. The related N₂ sorption isotherms along with the pore size distribution graphs are depicted in Figure 8. According to Figure 9-11, the N₂ sorption isotherms exhibit the typical type I curve with sharp N₂ adsorption at low partial pressures ($P/P_0 < 0.01$) and a H₂ hysteresis loop. In addition, the adsorbents have conical pores and crystals were composed of

the microporous structure. The specific surface area and pore volume of MIL-101(Cr) were calculated and summarized in Table 1. The results indicated that MIL-101(Cr) possessed the BET surface area ($2155\text{ m}^2\text{g}^{-1}$) and the Barrett–Joyner–Halenda (BJH) pore size distribution curve (b) derived from adsorption data of the isotherms indicates a main peak with average pore width of 2.1 for MIL-101(Cr). The result implies the microporous nature of samples and is in agreement with the results deduced from nitrogen adsorption-desorption isotherms and pore size distribution.

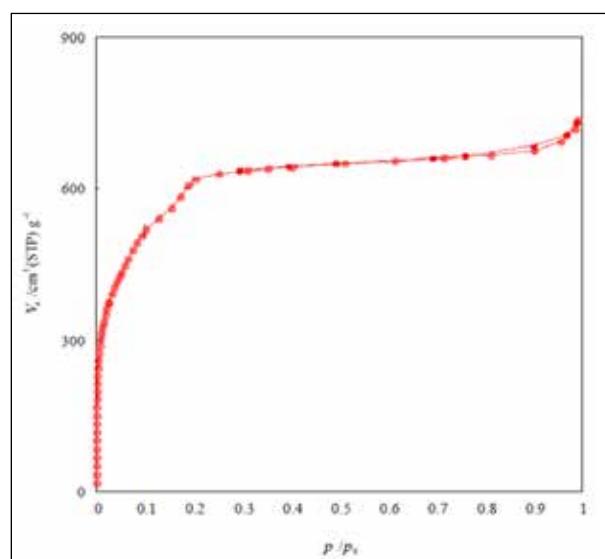


Fig. 8. Adsorption/desorption isotherm of Cr-MOF adsorbent with adsorptive of N₂ at 77oK

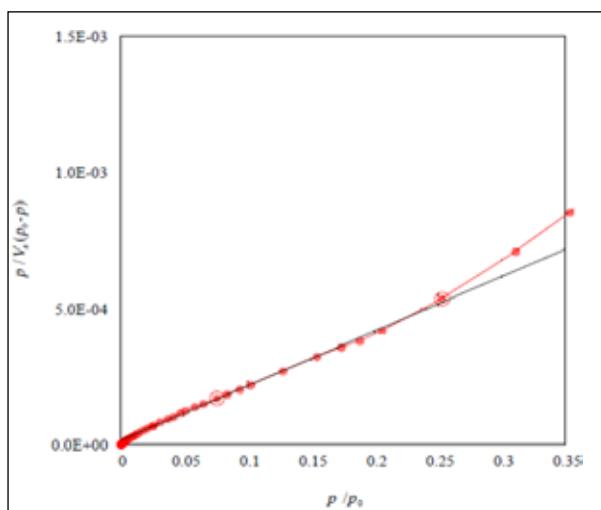


Fig. 9. BET plot of Cr-MOF adsorbent with adsorptive of N_2 at 77°K

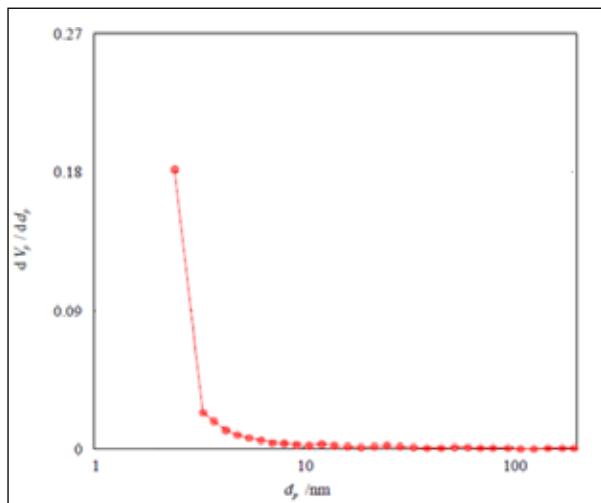


Fig. 10. BJH plot of Cr-MOF adsorbent for absorption branch at 77°K

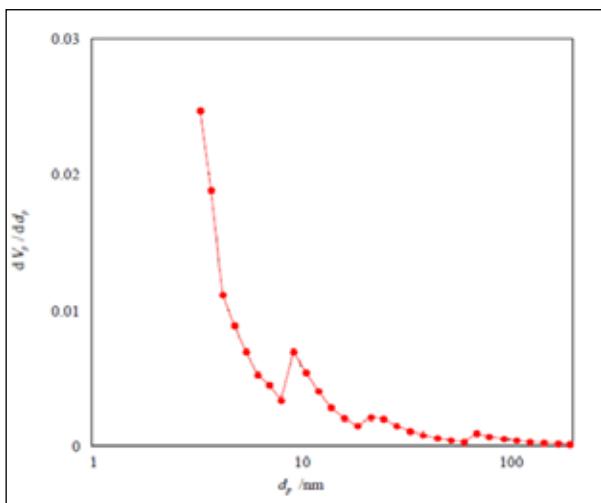


Fig. 11. BJH-plot of Cr-MOF adsorbent for desorption branch at 77°k

Table 1. Physical properties of Cr-MOF (MIL-101(Cr)) adsorbent

Product	BET Surface area ($\text{m}^2 \text{ g}^{-1}$)	Average pore width (nm)	V_m ($\text{cm}^3 \text{ g}^{-1}$)	Total pore volume ($\text{cm}^3 \text{ g}^{-1}$)
MOF	2155	2.1	495	1.13

3.6. Optimization

The DS- μ -SPE procedure based on Cr-MOF adsorbent was applied for speciation and preconcentration of nickel in water samples. For optimization many parameters such as amount of sorbent, volume of samples, shaking time, interference Ions and pH were optimized.

3.6.1. Amount of sorbent

For efficient extraction of nickel, the amount of MIL-101(Cr) nanoadsorbent was studied. For this purpose, the amounts of 5 - 50 mg of MIL-101(Cr) were used for nickel extraction by the DS- μ -SPE method. The results showed, the high recoveries between 95-105% were obtained in wastewater samples with 28 mg of MIL-101(Cr) for Ni extraction. Therefore, 30 mg of MOF was used as optimum mass for Ni extraction at $\text{pH}=8.5$ (Fig. 12).

3.6.2. Volume of samples

The sample volume for nickel extraction based on MOF was evaluated by DS- μ -SPE method. In this study the vary volume of wastewater samples between 5-100 mL was studied and optimized for $5\text{-}160 \mu\text{g L}^{-1}$ of nickel concentration. Based on results, the efficient extractions were achieved for 55 mL of wastewater samples. Therefore, 50 mL blood sample was selected as the optimal sample volume for further study (Fig. 13).

3.6.3. The Shaking time

The time of extraction depended on the sonication of MIL-101(Cr) nanoadsorbent in the wastewater samples, which was increased interaction between carboxyl group of MIL-101(Cr) with $\text{Ni}(\text{II})$ at

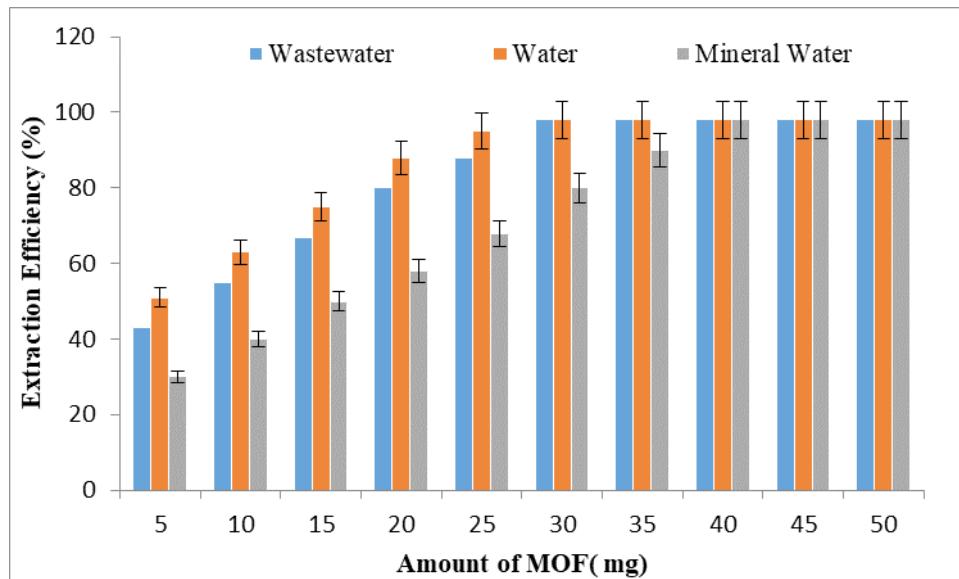


Fig. 12. The effect of MOF amount for Ni(II) extraction from water samples

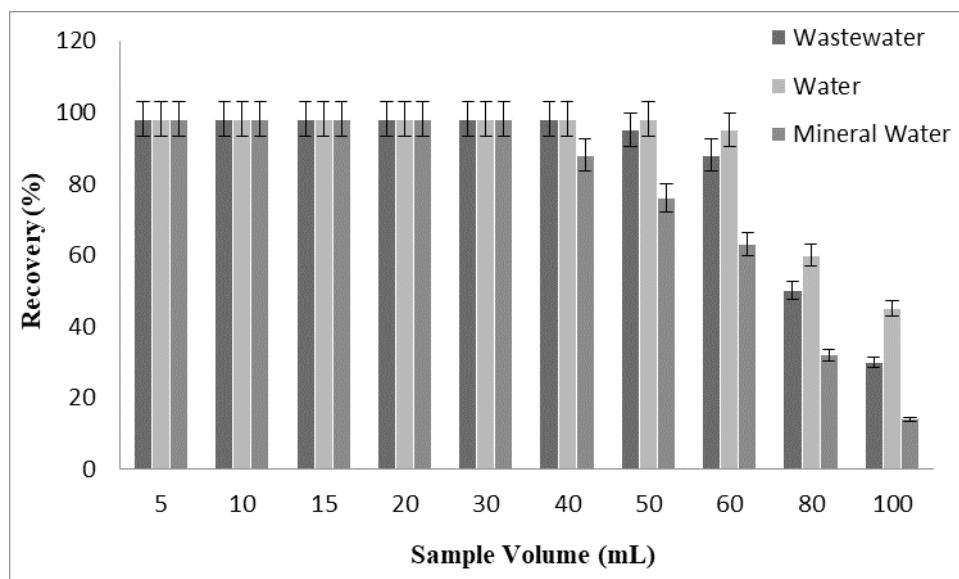


Fig.13. The effect of sample volume for Ni(II) extraction from water samples

pH=8.5. By dispersion of the MIL-101(Cr), the mass-transference and extraction was performed. So, the effect of shaking time was studied for 1-20 min. By results, the sonication of 5.0 min had efficient extraction for Ni ions in wastewater samples for 30 mg of sorbent. Therefore, 5.0 minute was used as the optimum shaking time by DS- μ -SPE procedure.

3.6.4. The pH effect

The pH has main role for nickel extraction by MIL-101(Cr) nanoadsorbent. So, the effect of different pH (2-11) for extraction of Ni (II) in wastewater samples were investigated by DS- μ -SPE procedure. The results demonstrated that the MIL-101(Cr) as MOF sorbent could be efficiently captured Ni (II) in pH of 8-8.5. Also, the recoveries

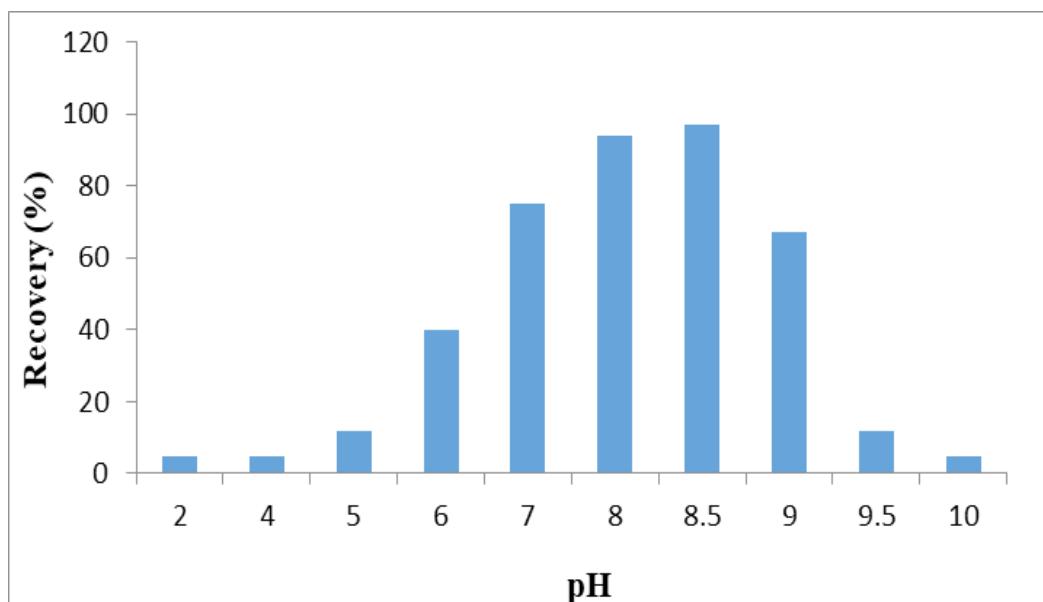


Fig. 14. The effect of pH for Ni(II) extraction based on MOF from water samples

for Ni (II) based on Cr-MOF were obtained more than 95% in pH 8.5 and the decreased at lower and higher pH ($7.5 > \text{pH} > 9$). So, pH=8.5 was selected for further experiments (Fig. 14). The mechanism of nickel extraction depend on the coordination of covalent bond of sorbent [$\text{Ni}^{2+} \rightarrow (\text{COO})^2-$] with the positively charged of nickel.

3.6.5. Validation

The validation of DS- μ -SPE procedure based on MOF was obtained by spiking samples. The

different concentration of nickel added to real samples as lower and upper ranges in wastewater and water samples. The results showed a good recovery between 95%-106% which was shown in Table 2. Also, the validation was confirmed with electrothermal atomic absorption spectrometry (ET-AAS) coupled with microwave digestion process (Table 3). Based on results in Tables 2 and 3, the nickel ions were efficiently extracted by Cr-MOF adsorbent in optimized conditions.

Table 2. The validation of DS- μ -SPE procedure based on MOF for nickel(II) extraction from water samples

Samples	Added ($\mu\text{g L}^{-1}$)	*Found ($\mu\text{g L}^{-1}$)	Recovery (%)
Water	-----	86.75 ± 3.55	-----
	100	185.63 ± 8.73	98.8
Tab water	-----	6.22 ± 0.34	-----
	5	11.53 ± 0.48	106.2
^a Wastewater	-----	144.62 ± 6.13	-----
	150	288.49 ± 13.26	95.9
Well water	-----	38.54 ± 2.05	-----
	40	77.68 ± 3.06	97.8

* Mean of three determinations \pm confidence interval ($P=0.95$, $n=8$)

^a Wastewater samples diluted with DW (1:1)

Table 3. Validation of methodology for nickel extraction from water samples by comparing to ET-AAS coupled with microwave digestion process

Samples	Added ($\mu\text{g L}^{-1}$)	* Found ETAAS ($\mu\text{g L}^{-1}$)	* Found AT-FAAS ($\mu\text{g L}^{-1}$)	Recovery ETAAS (%)	Recovery AT-FAAS (%)
Water 1	-----	10.56 \pm 0.43	10.18 \pm 0.52	-----	-----
	10	20.31 \pm 0.86	19.87 \pm 0.93	97.5	96.9
Water 2	-----	33.82 \pm 1.57	31.95 \pm 1.63	-----	-----
	30	62.98 \pm 2.78	60.88 \pm 3.02	97.2	96.4
Water 3	-----	75.21 \pm 3.26	73.93 \pm 3.39	-----	-----
	50	125.73 \pm 6.03	122.86 \pm 6.24	101.1	97.8
Water 4	-----	93.12 \pm 4.16	94.27 \pm 4.43	-----	-----
	100	191.62 \pm 8.94	194.77 \pm 9.24	98.5	100.5

* Mean of three determinations \pm confidence interval (P= 0.95, n=8)

Water1: Darband River,

Water 2: Hesarak River

Water 3: DarAbad River

Water 4: Zaferaniyeh and Velenjak River

3.6.6. Discussions

For metal determination with the DS- μ -SPE procedure based on Cr-MOF adsorbent, the effect of the main parameters, such as amount of sorbent, volume of samples, shaking time, interference Ions and pH were optimized thoroughly. Results this study revealed that MIL-101(Cr) nanoadsorbent was used as a novel sorbent for dispersive suspension micro solid phase extraction (DS- μ -SPE) nickel from environmental waters. In in study highly crystalline octahedral morphology of MIL-101(Cr) sample as reported MIL-101(Cr) structures in other studies was characterized by field emission scanning electron microscope (FE-SEM) analysis [64,65]. FT-IR characterization was conducted to detect the identity of the MIL-101(Cr) functional groups that the FT-IR results were in accordance to the previous reports [59, 64-66]. In this study, optimum mass of MOF, pH and time for Ni extraction 30 mg, 7.5 and 5 minute at pH: 7.5, 30 respectively were obtained in wastewater samples so that Behbahani, M., et al. this parameters for Modification of magnetized MCM-41 by pyridine groups for ultrasonic-assisted dispersive micro-solid-phase extraction of nickel ions 24 mg, 7.5 and 8 minute has been reported [67]. Based on results,

under optimized conditions detection limit, the linear range were achieved of $1.5 \mu\text{g L}^{-1}$, $5\text{-}160 \mu\text{g L}^{-1}$. In study safavi et al. linear range for 2-amino-cyclopentene-1-dithiocarboxylicacid absorbent and Cloud point extraction method linear in the range of $20\text{-}500 \mu\text{g L}^{-1}$ Obtained [22]. The findings of this study showed that MIL-101(Cr) as a valid procedure for extraction nickel in water samples can be used.

4. Conclusions

The MIL-101(Cr) as a MOF nanoadsorbent was synthesized and used for nickel extraction from water samples by DS- μ -SPE procedure. After extraction nickel ions with MOF adsorbent at pH=8.5, it was back-extracted from adsorbent and finally, the concentration was determined by AT-FAAS. The low LOD of $1.5 \mu\text{g L}^{-1}$ and the favorite linear range $5\text{-}160 \mu\text{g L}^{-1}$ was achieved. The high absorption capacity of 136.8 mg g^{-1} was obtained by MIL-101(Cr). The efficient, simple and fast extraction was achieved in low time by DS- μ -SPE procedure. The MIL-101(Cr) as solid-phase had high recovery more than 97% for Ni(II) extraction from waters without any chelating agents. Therefore, extraction of nickel based on

MIL-101(Cr) can be used as efficient procedure for determination and separation of Ni(II) in water samples by AT-FAAS

5. Acknowledgements

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